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Common Core Tennessee Book. Chapter 10 Vocabulary "The Mole". Terms in this set (12) Avogadro's Number. "The number  $6.0221367 \times 10^{23}$ , which is the number of representative particles in a mole, and can be rounded to three significant digits  $6.02 \times 10^{23}$ ". Empirical Formula.

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BJU Chemistry Chapter 10 Vocabulary. diffusion. effusion. Graham's law of effusion. permeability. the spontaneous, uniform spreading and mixing of particles cau.... a process in which gas particles pass through a tiny opening i.... the gas law that states that the rate of effusion is inversely....

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Chapter 10 & 11 Vocab CHEMISTRY. STUDY. PLAY. Kinetic molecular theory. a theory that explains that the behavior of physical systems depends on the combined actions of the molecules constituting the system. Ideal gas. an imaginary gas whose particles are infinitely small and do not interact with each other.

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AJScheick. Chemistry Chapter 10 liquids and solids. Johann van Helmont 1662. gas is derived from. vapor. vapor is formed if. used "gas" to describe the most energetic phase of matter. Greek "chaos" meaning original matter of the earth. a gas produced from a substance that is a solid or liquid unde....

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### **chemistry vocab chapter 10 liquids solids Flashcards and ...**

Chemistry Vocab Chapter 10. Avogadro's Number. mole. molar mass. Empirical Formula.  $6.02 \times 10^{23}$ ; counts extremely small particles. an SI base unit used to measure the quantity of matter. Mass in grams of one mole of any pure substance. Formula with smallest whole-number ratio of the elements.

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List of Important Chemistry Vocabulary Words. Share Flipboard Email Print It's much easier to learn chemistry if you understand the meaning of vocabulary words. There are a lot of them, so a good dictionary or glossary helps!. Andrey Prokhorov / Getty Images Science. Chemistry Basics Chemical Laws

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1405- chapter 10 Vocab (1).rtf - Chemistry 1405 Chapter 10 1... This preview shows page 1 - 2 out of 2 pages. Chemistry 1405 Chapter 10 1. Stoichiometry – calculate the amount of reactants required for a chemical reaction to be complete, and the amount of products made after the reaction comes to an end. 2. Use of a chemical equation to obtain conversion factors – a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles.

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### **1405- chapter 10 Vocab(1).rtf - Chemistry 1405 Chapter 10 ...**

Chapter 10 Vocabulary 1. Pressure-a measure of force exerted on a unit area. 2. Pascals-The SI unit of pressure 3. Standard atmospheric pressure- defined as 760 torr or, in SI units 101.326 kPa. 4. Atmosphere (atm)- a unit of pressure equal to 760 torr; 1 atm= 101.325 kPa. 5. Torr-a unit of pressure 1 torr= 1 mm Hg 6.

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