

## Chapter 19 Acids Bases Answers

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### Chapter 19 Acids Bases Answers

An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous solution. Bronsted-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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Answers Chapter 19 - Acids, Bases, and Salts - 19 Assessment - Page 684: 52 Answer The Arrhenius definition states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H<sup>+</sup>)

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Answer.  $H_2SO_4 + 2KOH \rightarrow 2H_2O + K_2SO_4$   $H_3PO_4 + 3Ca(OH)_2 \rightarrow 6H_2O + Ca_3(PO_4)_2$   $HNO_3 + Mg(OH)_2 \rightarrow 2H_2O + Mg(NO_3)_2$ . Work Step by Step. To write a complete balanced equation for the acid-base reaction, take the cation from the base and the anion from the acid to form the salt. The other product is water.

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### Acids And Bases Chapter 19 Answers

1-. concentration to be greater than in pure water, and the ammonia solution is basic. Acids and bases come in pairs. A " conjugate base. is the remainder of the original acid, after it donates its hydrogen ion. A " conjugate acid. is the particle formed when the original base gains a hydrogen ion.

### Chapter 19 Acids, Bases, and Salts

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### Science / Chapter 19 - acids, bases, and salts (handouts)

Acid and Base Worksheet - Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a)  $HNO_3 + OH^- \rightarrow H_2O + NO_3^-$ .  $HNO_3$  and  $NO_3^-$  make one pair  $OH^-$  and  $H_2O$  make the other. b)  $CH_3NH_2 + H_2O \rightarrow CH_3NH_3^+ + OH^-$

### Acid and Base Worksheet - Answers - Chemistry Made Easy

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Chapter 19 - Acids & Bases Chapter 19 covers properties of acids and bases, the different definitions,  $[H^+]$  and  $[OH^-]$  and pH and pOH,  $K_a$  and  $K_b$  and their use when dealing with weak acids and bases and buffers.

**Chapter 19 Acids Bases Answer Key**

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