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In the atoms of oxygen the valence electrons are 6 (i.e., electrons in the outermost shell). To fill the orbit, 2 electrons are required. In the atom of silicon, the valence electrons are 4. To fill this orbit 4 electrons are required. Hence, the combining capacity of oxygen is 2 and of silicon is 4.

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Molecular Structure includes all the important topics with detailed explanation that aims to help students to understand the concepts better. Students who are preparing for their Class 11 exams must go through NCERT Solutions for Class 11 Chemistry Chapter 4 Chemical Bonding and Molecular Structure.

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NCERT Solutions for Class 9 Science Chapter 4 - Structure of the Atom. In this chapter, we get enlightened about the discovery of electrons and protons that is credited to J.J. Thomson and E. Goldstein, respectively. As per the theory by J.J. Thomson, electrons are implanted in the positive sphere.

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RBSE Class 12 Chemistry Chapter 4 Long Answer Type Question. Question 1. The decomposition of dimethyl ether leads to

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the formation of CH_4 , H_2 and CO and the reaction rate is given by, $\text{Rate} = k [\text{P} \text{CH}_3 \text{OCH}_3]^{3/2}$. If the pressure is measured in bar and time in minutes, then what are the units of the rate and rate constant? Answer: $\text{CH}_2 \text{OCH}_3 \dots$

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