

## Section 22 1 Nuclear Chemistry Answers

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### CHAPTER 22 Nuclear Chemistry

Chemistry NUCLEAR CHEMISTRY Section 22-1: The Nucleus Objectives 1. Explain what nucleons are. 2. Explain what a nuclide is, and describe the different ways it can be written. 3. Define nuclear

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binding energy. 4. Explain the relationship between nucleon number and stability of nuclei. Chapter 22 - Nuclear

### Chapter 22 Review Nuclear Chemistry Mixed

Chapter 22: Nuclear Chemistry Section 22-1: The Nucleus • Atomic nuclei= protons and neutrons (together are nucleons) o Nuclide= an atom—identified by # of protons/neutrons in nucleus Mass Defect and Nuclear Stability • Mass defect= difference between mass of an atom and sum of the masses of protons/neutrons/electrons o Caused by conversion of mass to energy when nucleus forms Nuclear Binding Energy •  $E=mc^2$ — mass can be converted to energy+energy can be converted to mass ...

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prefer chapter 22 nuclear chemistry section 22 1 the nucleus o atomic nuclei protons and neutrons

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together are nucleons o nuclide an atom identified by of protons neutrons in nucleus mass defect and nuclear stability o mass defect difference between mass of an atom and sum of the masses of protons

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By the end of this section, you will be able to: Describe nuclear structure in terms of protons, neutrons, and electrons. Calculate mass defect and binding energy for nuclei. Explain trends in the relative stability of nuclei. Nuclear chemistry is the study of reactions that involve changes in nuclear structure.

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other having 39 neutrons: charge applicable - X-element symbol (see periodic table) Atomic Number=2 A cation with a charge of +3 containing 13 protons and 14 neutrons: An anion with a charge of -1 containing 53 ...

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